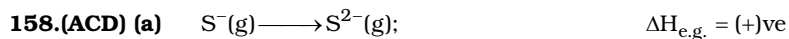


**Daily Tutorial Sheet-15**

**Level-3**

**157.(C)** If Aufbau rule is not followed then, 19th electron in K enters in 3d sub-shell, not in 4s.

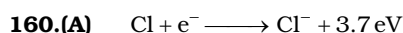


**159.(ABD)(A)** Cations are small than parental atom due to less no. of  $e^{-}$

**(B)** Size decreases due to presence of less  $e^{-}$  as compared to proton and hence attraction towards proton increases.

**(C)** Incorrect  $\rightarrow$  Correct order is :  $Co \approx Ni < Cu < Zn$

**(D)** Down the group size increases. So, order is  $K^{+} > Mg^{2+} > Al^{3+} > Li^{+}$



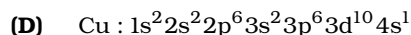
Energy released for the conversion of 2.0 g of gaseous chlorine to chloride ion

$$= \frac{3.7 \times 23.06 \times 2}{35.5} = 4.8 \text{ kcal}$$

**161. (A)** 3s-orbital cannot be filled before complete filling of 2p-orbital. So, this configuration is incorrectly written.

**(B)** Atomic number of boron is 5, so its electron configuration would be  $1s^2 2s^2 2p^1$ , the configuration written above is that of fluorine.

**(C)** Atomic number of chlorine is 17, so its electronic configuration would be  $1s^2 2s^2 2p^6 3s^2 3p^5$ . The configuration mentioned above is that of Argon.



**162.(ACD)** On Pauling's scale electronegativities of H (2.1), Te (2.1) and P (2.1) are similar but the electronegativity of S (2.5) is different from the other three elements.